



**Cambridge International Examinations**  
Cambridge International General Certificate of Secondary Education

CANDIDATE  
NAME

CENTRE  
NUMBER

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CANDIDATE  
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**CHEMISTRY**

**0620/33**

Paper 3 Theory (Core)

**October/November 2016**

**1 hour 15 minutes**

Candidates answer on the Question Paper.

No Additional Materials are required.

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 16.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **16** printed pages.



(b) The table gives some information about the properties of four metals.

metal	density in g/cm <sup>3</sup>	relative strength	resistance to corrosion	relative electrical conductivity	melting point/°C
chromium	7.2	8	very good	8	1857
copper	8.9	30	good	60	1283
iron	7.9	21	poor	10	1535
titanium	4.5	23	very good	2	1660

Which **one** of these metals is most suitable for making the frame of an aircraft?  
Explain your answer using information from the table.

.....

.....

.....

..... [3]

[Total: 8]

- 2 A scientist analysed the substances present in a 1 dm<sup>3</sup> sample of river water in an agricultural area. The table shows the mass of each ion dissolved in the 1 dm<sup>3</sup> sample.

name of ion	formula of ion	mass/g
calcium	Ca <sup>2+</sup>	1.2
chloride	Cl <sup>-</sup>	0.1
hydrogencarbonate	HCO <sub>3</sub> <sup>-</sup>	1.0
magnesium	Mg <sup>2+</sup>	0.5
nitrate	NO <sub>3</sub> <sup>-</sup>	1.0
sodium	Na <sup>+</sup>	
	SO <sub>4</sub> <sup>2-</sup>	0.5
phosphate	PO <sub>4</sub> <sup>3-</sup>	1.2
	Total	6.0

- (a) (i) Which negative ion has the highest concentration, in g/dm<sup>3</sup>, in this sample of water?

..... [1]

- (ii) Give the name of the ion with the formula SO<sub>4</sub><sup>2-</sup>.

..... [1]

- (iii) Calculate the mass of sodium ions in 1 dm<sup>3</sup> of this river water.

..... [1]

- (b) Describe a test for nitrate ions.

test .....

.....

result .....

[3]

(c) The sample of river water also contains insoluble materials such as clay and the remains of dead animals and plants.

(i) What method could be used to separate insoluble materials from river water?

..... [1]

(ii) Some of the remains of dead animals and plants contain food materials.

Which **two** of the following substances are constituents of food?  
Tick **two** boxes.

alkane	<input type="checkbox"/>
carbohydrate	<input type="checkbox"/>
graphite	<input type="checkbox"/>
protein	<input type="checkbox"/>

[1]

(iii) Particles of clay suspended in river water show Brownian motion.

Describe the movement of these particles.

..... [1]

(d) Most of the nitrate ions in river water come from fertilisers.

(i) Explain why farmers use fertilisers.

.....  
..... [2]

(ii) Ammonium nitrate is a fertiliser.  
Ammonium nitrate reacts with calcium hydroxide.

ammonium nitrate + calcium hydroxide → calcium nitrate + ammonia + water

Explain why adding calcium hydroxide to the soil at the same time as nitrate fertilisers results in loss of nitrogen from the soil.

.....  
..... [2]

[Total: 13]

3 Ethanol can be manufactured by fermentation and from ethene.

(a) Describe the manufacture of ethanol by fermentation **and** from ethene.  
In your answer include

- the essential conditions required for each reaction,
- **one** or more relevant word equations.

.....

.....

.....

.....

.....

.....

..... [5]

(b) The table shows some properties of different alcohols.

alcohol	formula	melting point/°C	boiling point/°C	relative viscosity
methanol	CH <sub>4</sub> O	-94	65	0.54
ethanol	C <sub>2</sub> H <sub>6</sub> O	-117	79	1.08
propanol	C <sub>3</sub> H <sub>8</sub> O	-126	98	1.94
butanol	C <sub>4</sub> H <sub>10</sub> O	-89	117	2.54
pentanol	C <sub>5</sub> H <sub>12</sub> O	-79		3.47

(i) Deduce the state of methanol at room temperature.  
Explain your answer.

.....

..... [2]

(ii) Predict the boiling point of pentanol.

..... [1]

(iii) Describe how the relative viscosity changes with the number of carbon atoms in the alcohol.

..... [1]

(c) (i) Draw the structure of ethanol. Show all of the atoms and all of the bonds.

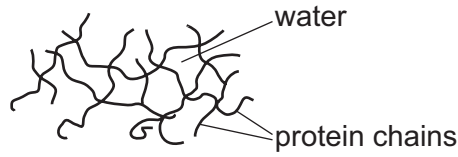
[2]

(ii) Give **one** major use of ethanol.

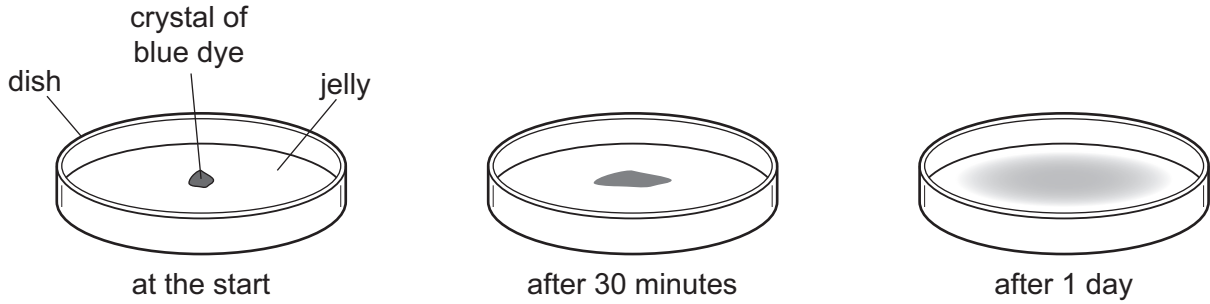
..... [1]

[Total: 12]

4 Jelly is a mixture of water and protein chains.



(a) A crystal of blue dye was placed on top of some jelly.  
 After 30 minutes some of the blue colour could be seen in the jelly.  
 After 1 day the blue colour had spread out further into the jelly.



Use the kinetic particle model of matter to explain these observations.

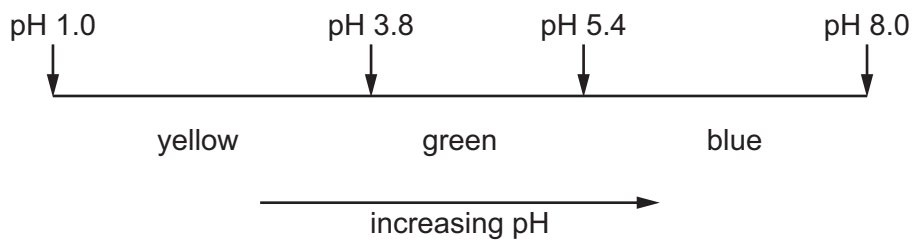
.....

.....

.....

..... [3]

(b) The diagram shows the colour changes of the indicator bromocresol green at different pH values.



Predict the colour of bromocresol green

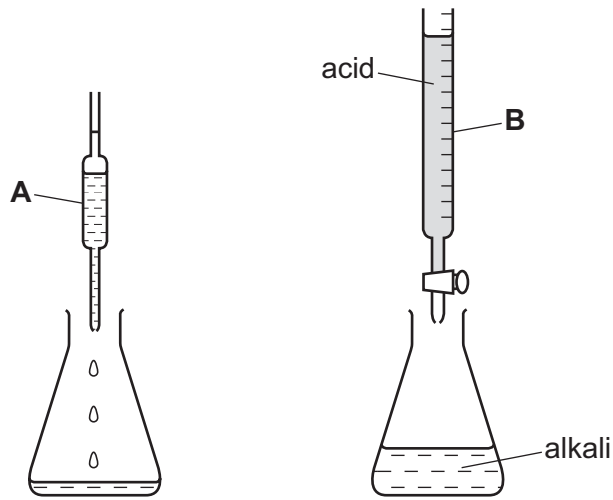
in pure water, .....

in a strongly acidic solution. ....

[2]



(c) The concentration of an alkali can be found by titrating it with an acid using the apparatus shown.



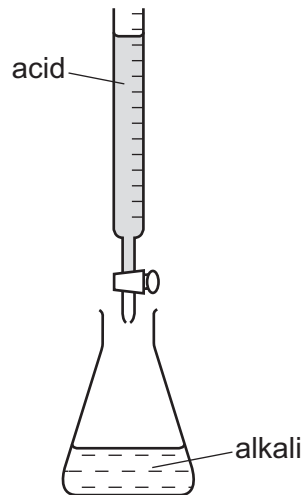
(i) State the names of the pieces of glassware labelled **A** and **B**.

**A** .....

**B** .....

[2]

(ii) Describe how you would carry out a titration using the apparatus shown.



.....

.....

.....

.....

.....

.....

.....

.....

.....

.....

[3]

[Total: 10]

- 5 Lime (calcium oxide) is made by heating limestone (calcium carbonate).



- (a) (i) Is this reaction exothermic or endothermic?  
Explain your answer.

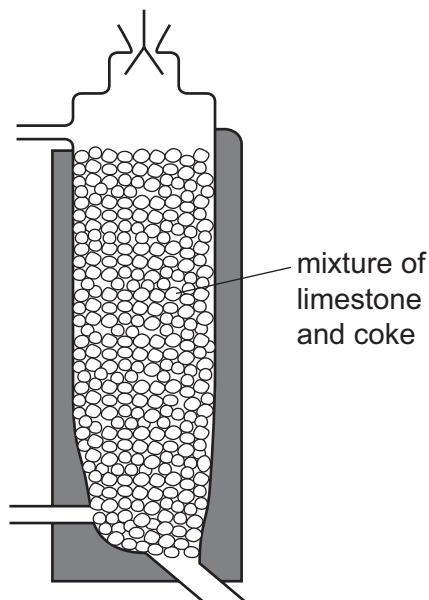
.....  
..... [1]

- (ii) The reaction is reversible.

What information in the equation shows that this reaction is reversible?

..... [1]

- (b) The diagram shows a furnace for making lime.



- (i) On the diagram, write

- the letter **C** to show where the waste gases exit the furnace,
- the letter **L** to show where the lime is removed from the furnace.

[2]

- (ii) Suggest a reason for adding coke (carbon) to the furnace.

..... [1]

- (c) Explain why farmers use lime to treat acidic soils.

.....  
..... [2]

(d) Limestone is used to manufacture cement. The limestone is mixed with clay and heated to 1500 °C. It is then mixed with calcium sulfate and crushed.

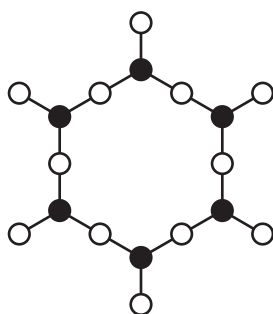
(i) Describe the test for sulfate ions.

test .....

result .....

[2]

(ii) Concrete is a mixture of cement, silicates and water. Part of the structure of a silicate is shown.



key

● = silicon atom

○ = oxygen atom

Deduce the formula for this silicate.

..... [1]

- (e) Concrete contains small amounts of calcium oxide.  
This can react with rainwater to form calcium hydroxide.

- (i) Calcium hydroxide is strongly alkaline.

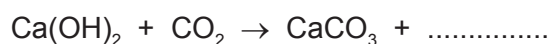
What is the most likely pH of a strongly alkaline solution?  
Draw a ring around the correct answer.

pH 2                  pH 6                  pH 7                  pH 12

[1]

- (ii) The calcium hydroxide on the surface of a piece of concrete reacts with carbon dioxide in the air.

Complete the chemical equation for this reaction.



[1]

- (iii) Limewater is an aqueous solution of calcium hydroxide. A teacher left an open beaker of limewater in the laboratory.  
After a week, the solution in the beaker was pH 7 and a white precipitate was observed.

Use the information in (e)(i) and (e)(ii) to help you explain these observations.

.....  
.....  
.....  
.....  
.....  
.....  
..... [3]

[Total: 15]

6 The Periodic Table is a method of classifying elements.

(a) (i) In what order are the elements arranged in the Periodic Table?

..... [1]

(ii) How does the character of the elements change from left to right across a period?

..... [1]

(iii) Describe **two** trends in the properties of the elements going down Group I.

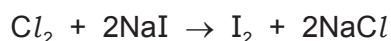
.....

.....

..... [2]

(b) The halogens are a group of elements with diatomic molecules.

(i) Chlorine reacts with an aqueous solution of sodium iodide.



What colour change would be observed in the solution?

from ..... to ..... [2]

(ii) Astatine, At<sub>2</sub>, is a halogen.

Suggest why astatine does **not** react with aqueous potassium iodide.

..... [1]

(c) Chlorine reacts with hydrogen to form hydrogen chloride.

(i) Complete the chemical equation for this reaction.



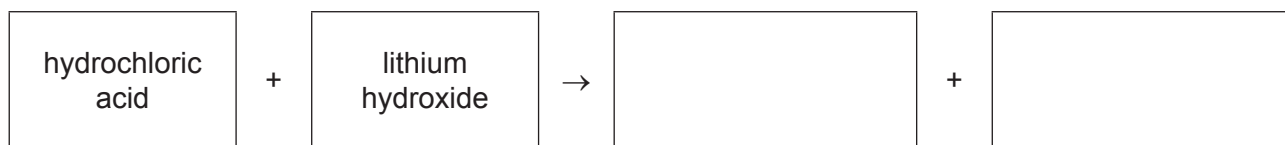
[2]

(ii) Draw a diagram to show the electronic structure of a molecule of hydrogen chloride. Show only the outer shell electrons.

[2]

(iii) Hydrochloric acid reacts with lithium hydroxide.

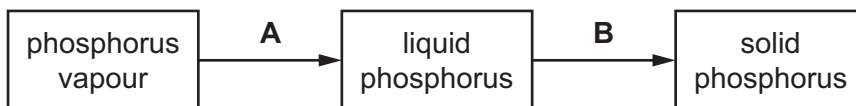
Complete the word equation for this reaction.



[2]

[Total: 13]

- 7 The diagram shows the changes of state when phosphorus is cooled slowly to room temperature.



- (a) Give the names of the changes of state labelled **A** and **B**.

**A** .....

**B** .....

[2]

- (b) Describe the arrangement and motion of the particles in solid phosphorus.

arrangement .....

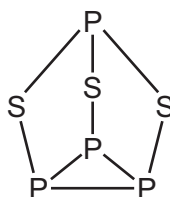
motion .....

[2]

- (c) Is phosphorus(V) oxide an acidic oxide or basic oxide?  
Explain your answer.

..... [1]

- (d) Phosphorus sulfide is a covalent molecule.



Predict **two** properties of phosphorus sulfide.

.....

..... [2]

- (e) Many metal ores contain sulfides.  
When zinc sulfide is heated in air the following reaction takes place.



Explain why this reaction may be harmful to the environment.

.....

..... [2]

[Total: 9]

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## The Periodic Table of Elements

Group																																			
I	II	III										IV	V	VI	VII	VIII																			
3 <b>Li</b> lithium 7	4 <b>Be</b> beryllium 9	<div style="border: 1px solid black; padding: 5px; text-align: center;"> <b>Key</b>            atomic number            atomic symbol            name            relative atomic mass         </div>										5 <b>B</b> boron 11	6 <b>C</b> carbon 12	7 <b>N</b> nitrogen 14	8 <b>O</b> oxygen 16	9 <b>F</b> fluorine 19	10 <b>Ne</b> neon 20																		
11 <b>Na</b> sodium 23	12 <b>Mg</b> magnesium 24	13 <b>Al</b> aluminium 27	14 <b>Si</b> silicon 28	15 <b>P</b> phosphorus 31	16 <b>S</b> sulfur 32	17 <b>Cl</b> chlorine 35.5	18 <b>Ar</b> argon 40	19 <b>K</b> potassium 39	20 <b>Ca</b> calcium 40	21 <b>Sc</b> scandium 45	22 <b>Ti</b> titanium 48	23 <b>V</b> vanadium 51	24 <b>Cr</b> chromium 52	25 <b>Mn</b> manganese 55	26 <b>Fe</b> iron 56	27 <b>Co</b> cobalt 59	28 <b>Ni</b> nickel 59	29 <b>Cu</b> copper 64	30 <b>Zn</b> zinc 65	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84										
37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	39 <b>Y</b> yttrium 89	40 <b>Zr</b> zirconium 91	41 <b>Nb</b> niobium 93	42 <b>Mo</b> molybdenum 96	43 <b>Tc</b> technetium —	44 <b>Ru</b> ruthenium 101	45 <b>Rh</b> rhodium 103	46 <b>Pd</b> palladium 106	47 <b>Ag</b> silver 108	48 <b>Cd</b> cadmium 112	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	53 <b>I</b> iodine 127	54 <b>Xe</b> xenon 131	55 <b>Cs</b> caesium 133	56 <b>Ba</b> barium 137	57–71 lanthanoids	72 <b>Hf</b> hafnium 178	73 <b>Ta</b> tantalum 181	74 <b>W</b> tungsten 184	75 <b>Re</b> rhenium 186	76 <b>Os</b> osmium 190	77 <b>Ir</b> iridium 192	78 <b>Pt</b> platinum 195	79 <b>Au</b> gold 197	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —
87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	89–103 actinoids	104 <b>Rf</b> rutherfordium —	105 <b>Db</b> dubnium —	106 <b>Sg</b> seaborgium —	107 <b>Bh</b> bohrium —	108 <b>Hs</b> hassium —	109 <b>Mt</b> meitnerium —	110 <b>Ds</b> darmstadtium —	111 <b>Rg</b> roentgenium —	112 <b>Cn</b> copernicium —	113 <b>Nh</b> nihonium —	114 <b>Fl</b> flerovium —	115 <b>Mc</b> moscovium —	116 <b>Lv</b> livermorium —	117 <b>Ts</b> tennessine —	118 <b>Og</b> oganesson —																		

lanthanoids

57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	59 <b>Pr</b> praseodymium 141	60 <b>Nd</b> neodymium 144	61 <b>Pm</b> promethium —	62 <b>Sm</b> samarium 150	63 <b>Eu</b> europium 152	64 <b>Gd</b> gadolinium 157	65 <b>Tb</b> terbium 159	66 <b>Dy</b> dysprosium 163	67 <b>Ho</b> holmium 165	68 <b>Er</b> erbium 167	69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	71 <b>Lu</b> lutetium 175
89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —

actinoids

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).