



# Cambridge IGCSE™

**CHEMISTRY**

**0620/21**

Paper 2 Multiple Choice (Extended)

**October/November 2020**

**45 minutes**

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet  
Soft clean eraser  
Soft pencil (type B or HB is recommended)

## INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A, B, C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

## INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark. A mark will not be deducted for a wrong answer.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

This document has **16** pages. Blank pages are indicated.



1 Which gas has the slowest rate of diffusion?

- A  $\text{H}_2$                       B  $\text{NH}_3$                       C  $\text{CH}_4$                       D  $\text{CO}_2$

2 A mixture of colourless amino acids is separated using chromatography.

The solvent used is propanol.

The chromatogram is sprayed with a locating agent.

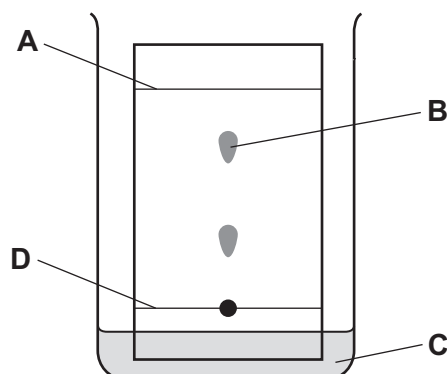
Which row describes the purpose of the propanol and the locating agent?

	purpose of propanol	purpose of locating agent
<b>A</b>	to make the individual amino acids visible	to prevent the amino acids moving any further
<b>B</b>	to move the amino acids up the chromatography paper	to make the individual amino acids visible
<b>C</b>	to move the amino acids up the chromatography paper	to prevent the amino acids moving any further
<b>D</b>	to prevent the amino acids moving too far up the paper	to make the individual amino acids visible

3 Which piece of apparatus can only measure a single fixed volume?

- A  $250\text{ cm}^3$  beaker  
 B  $50\text{ cm}^3$  burette  
 C  $100\text{ cm}^3$  measuring cylinder  
 D  $25\text{ cm}^3$  pipette

4 In the chromatography experiment shown, which label represents the solvent front?



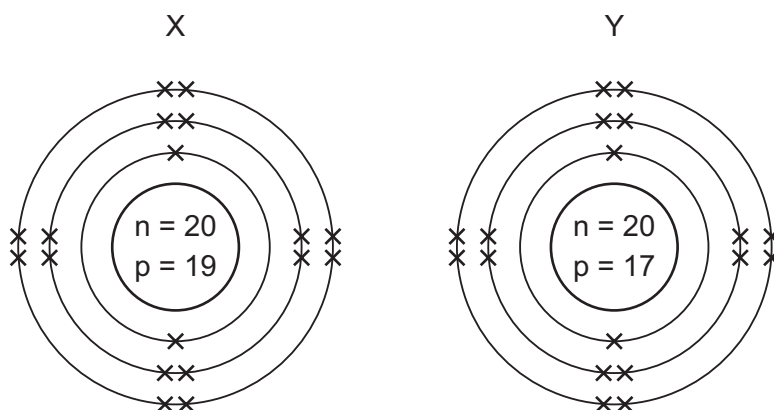
5 The atomic structure of four particles are shown.

	electrons	protons	neutrons
P	18	17	18
Q	18	17	20
R	17	17	18
S	17	17	20

Which particles have the same chemical properties?

- A** P and R only    **B** P and S    **C** P, Q and R    **D** R and S

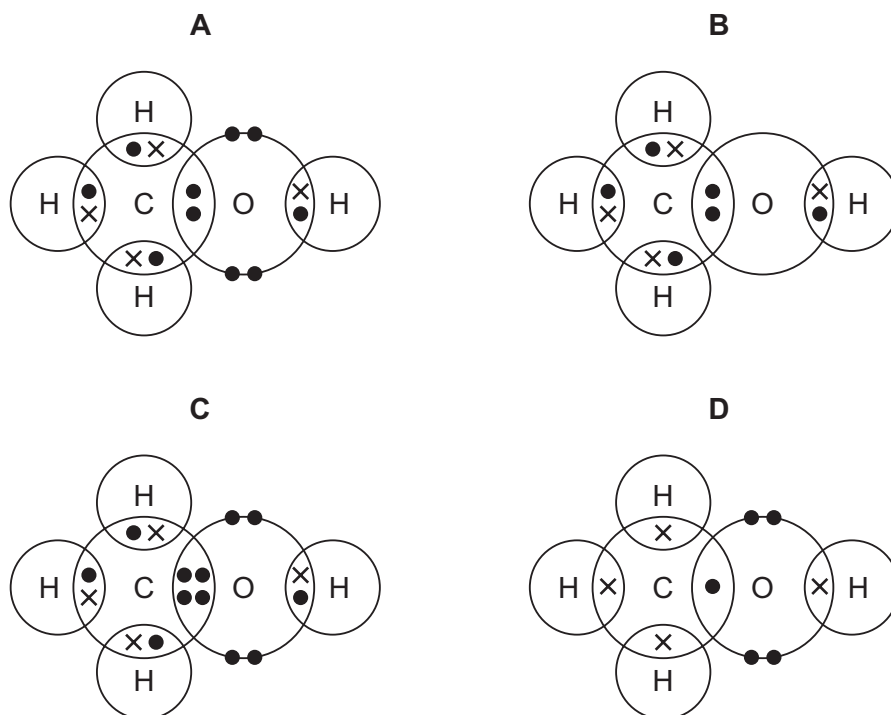
6 The arrangements of the electrons in two ions formed from elements X and Y are shown.



Which equation represents the reaction between elements X and Y?

- A**  $X_2 + 2Y \rightarrow 2X^+ + 2Y^-$   
**B**  $X_2 + 2Y \rightarrow 2X^- + 2Y^+$   
**C**  $2X + Y_2 \rightarrow 2X^+ + 2Y^-$   
**D**  $2X + Y_2 \rightarrow 2X^- + 2Y^+$

7 Which diagram shows the outer shell electron arrangement in a molecule of methanol,  $\text{CH}_3\text{OH}$ ?



8 Which statement about silicon dioxide,  $\text{SiO}_2$ , is correct?

- A** It conducts electricity because it contains free electrons.
- B** It is a macromolecule with four oxygen atoms bonded to each silicon atom.
- C** It is a simple covalent molecule.
- D** Its structure is similar to graphite.

9 Rubidium is in Group I of the Periodic Table and bromine is in Group VII.

Rubidium reacts with bromine to form an ionic compound.

Which row shows the electron change taking place for rubidium and the correct formula of the rubidium ion?

	electron change	formula of ion formed
<b>A</b>	electron gained	$\text{Rb}^+$
<b>B</b>	electron gained	$\text{Rb}^-$
<b>C</b>	electron lost	$\text{Rb}^+$
<b>D</b>	electron lost	$\text{Rb}^-$

10 Which statement explains why graphite is used as a lubricant?

- A All bonds between the atoms are weak.
- B It conducts electricity.
- C It has a low melting point.
- D Layers in the structure can slide over each other.

11 Sodium carbonate reacts with sulfuric acid to form carbon dioxide, water and a sodium salt.

An incomplete equation for the reaction is shown.



What is the formula of the sodium salt?

- A  $\text{Na}_2(\text{SO}_4)_2$       B  $\text{Na}(\text{SO}_4)_2$       C  $\text{Na}_2\text{SO}_4$       D  $\text{NaSO}_4$

12 The relative atomic mass of chlorine is 35.5.

When calculating relative atomic mass, which particle is the mass of a chlorine atom compared to?

- A a neutron
- B a proton
- C an atom of carbon-12
- D an atom of hydrogen-1

13 What is the empirical formula of an oxide of iron, formed by reacting 2.24 g of iron with 0.96 g of oxygen?

- A FeO                  B Fe<sub>2</sub>O                  C Fe<sub>2</sub>O<sub>3</sub>                  D Fe<sub>3</sub>O<sub>4</sub>

14 Which reaction takes place at the cathode during the electrolysis of molten nickel(II) chloride?

- A  $\text{Cl}_2 + 2\text{e}^- \rightarrow 2\text{Cl}^-$
- B  $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$
- C  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$
- D  $\text{Ni}^{2+} + 2\text{e}^- \rightarrow \text{Ni}$

15 Sodium nitrate is added to water in a beaker and stirred until it dissolves.

At the end of the experiment, the beaker feels cold.

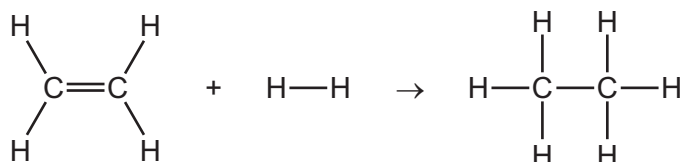
Which row describes the reaction?

	temperature of solution	type of reaction
<b>A</b>	decreases	endothermic
<b>B</b>	decreases	exothermic
<b>C</b>	increases	endothermic
<b>D</b>	increases	exothermic

16 Which substance does **not** require oxygen in order to produce energy?

- A** coal
- B** hydrogen
- C** natural gas
- D**  $^{235}\text{U}$

17 Ethene reacts with hydrogen to form ethane.



The bond energies are shown in the table.

bond	bond energy in kJ/mol
C–C	+350
C–H	+410
H–H	+436
C=C	+614

What is the energy change for the reaction?

- A** –290 kJ/mol
- B** –120 kJ/mol
- C** +120 kJ/mol
- D** +290 kJ/mol

18 A sign displayed in a flour mill is shown.



Which statement explains why there is a danger of explosion in a flour mill?

- A Flour burns very quickly because it is a fine powder.
- B Flour is a catalyst for combustion.
- C Flour mills get hot and speed up the rate of combustion.
- D The combustion of flour is exothermic.

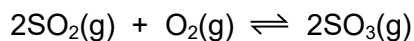
19 A student investigates the effect of concentration on the rate of reaction between calcium carbonate and hydrochloric acid. He follows the method shown.

- Place 1 g of calcium carbonate in a conical flask.
- Add excess hydrochloric acid.
- Let the reaction continue until no more gas is made.
- Repeat the experiment with different concentrations of hydrochloric acid.

Which essential step has been left out of the method if he is to work out the rate of the reaction?

- A heating the reaction mixture
- B placing a bung in the flask
- C timing the reaction
- D using a catalyst

20 The reaction between sulfur dioxide and oxygen is shown.



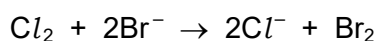
The reaction is exothermic.

Which of the changes shifts the position of equilibrium to the right?

- 1 Increase the concentration of oxygen.
- 2 Increase the pressure.
- 3 Increase the temperature.

**A** 1, 2 and 3      **B** 1 and 2 only      **C** 1 and 3 only      **D** 2 only

21 The reaction between chlorine and bromide ions is a redox reaction.



What is the change in oxidation state of the reducing agent in this reaction?

**A** -2 to 0      **B** -1 to 0      **C** 0 to -1      **D** 0 to +1

22 What is a characteristic of acids?

- A** Acids turn methyl orange indicator yellow.
- B** Acids have a high pH value.
- C** Acids react with ammonium salts to give ammonia gas.
- D** Acids react with carbonates to produce salts.

23 Zinc oxide is an amphoteric oxide.

Which row describes the reactions of zinc oxide?

	reaction with alkalis	reaction with acids
<b>A</b>	✓	✓
<b>B</b>	✓	x
<b>C</b>	x	✓
<b>D</b>	x	x





27 A flammable gas needs to be removed from a tank at an industrial plant.

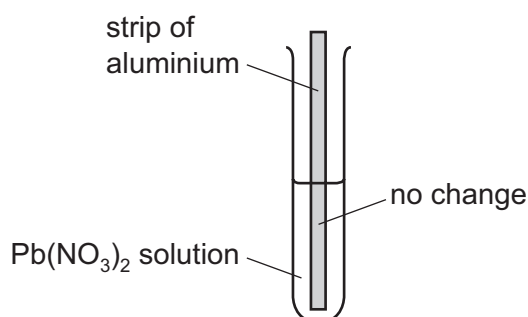
For safety reasons, an inert gas is used.

Which gas is suitable?

- A argon
- B hydrogen
- C methane
- D oxygen

28 A strip of aluminium is placed into a test-tube containing aqueous lead(II) nitrate and left for several minutes.

Aluminium is higher than lead in the reactivity series.



Which statement explains why lead is **not** displaced by this strip of aluminium?

- A A thin insoluble layer of aluminium nitrate forms on the aluminium.
- B Nitrate ions are reduced in aqueous solution.
- C The ionic bonds between lead and nitrate ions are too strong.
- D There is an unreactive oxide layer on the aluminium.

29 Which statements about the metal zinc are correct?

- 1 It is extracted from the ore bauxite.
- 2 It is used to galvanise steel.
- 3 It is used to make the alloy brass.
- 4 It reacts with dilute hydrochloric acid to produce hydrogen gas.

- A 2 and 3 only    B 1, 2 and 4    C 1, 3 and 4    D 2, 3 and 4

30 What is the symbol of the metal used in the manufacture of aircraft because of its strength and low density?

- A Al                      B Cu                      C Fe                      D Zn

31 Ammonia is manufactured using the Haber process.

Which statement about this process is correct?

- A The catalyst used for this reaction is vanadium pentoxide.
- B The hydrogen used is extracted from air.
- C Using a high pressure increases the yield of ammonia.
- D Using a high temperature increases the yield of ammonia.

32 Iron can be protected from rusting by attaching a piece of a more reactive metal, e.g. magnesium, to the iron.

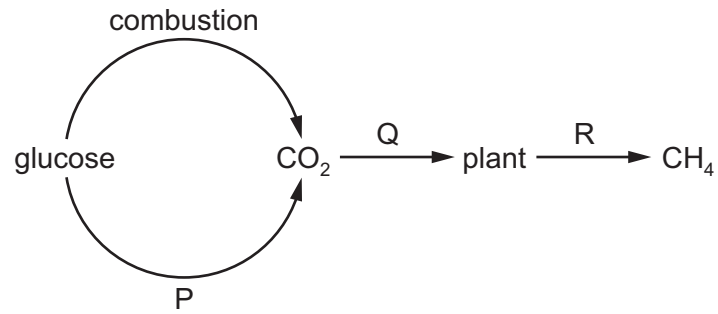
Which equation represents the reaction that takes place?

- A  $\text{Fe(s)} \rightarrow \text{Fe}^{2+}(\text{aq}) + 2\text{e}^{-}$
- B  $\text{Fe}^{2+}(\text{aq}) + 2\text{e}^{-} \rightarrow \text{Fe(s)}$
- C  $\text{Mg(s)} \rightarrow \text{Mg}^{2+}(\text{aq}) + 2\text{e}^{-}$
- D  $\text{Mg}^{2+}(\text{aq}) + 2\text{e}^{-} \rightarrow \text{Mg(s)}$

33 Which row describes two uses of sulfur dioxide?

	use 1	use 2
A	bleaching paper pulp	neutralising acidic industrial waste
B	bleaching paper pulp	preserving food and drink
C	extracting iron from hematite	neutralising acidic industrial waste
D	extracting iron from hematite	preserving food and drink

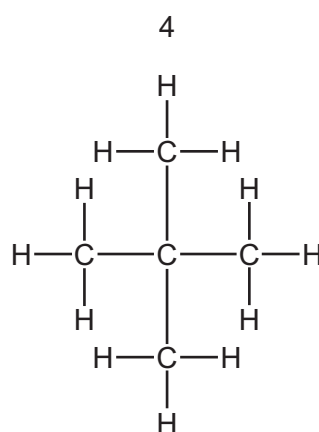
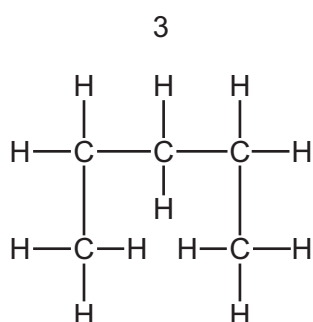
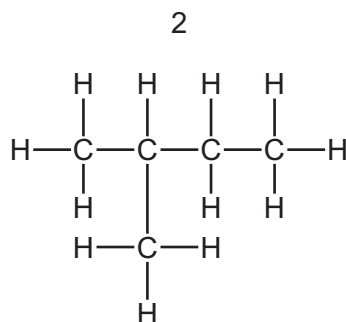
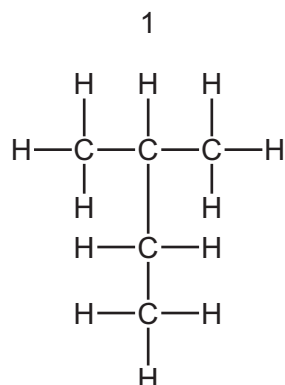
34 Part of the carbon cycle is shown.



What are processes P, Q and R?

	P	Q	R
<b>A</b>	decomposition	respiration	photosynthesis
<b>B</b>	respiration	photosynthesis	decomposition
<b>C</b>	respiration	decomposition	photosynthesis
<b>D</b>	photosynthesis	respiration	decomposition

35 The structures of four organic molecules are shown.



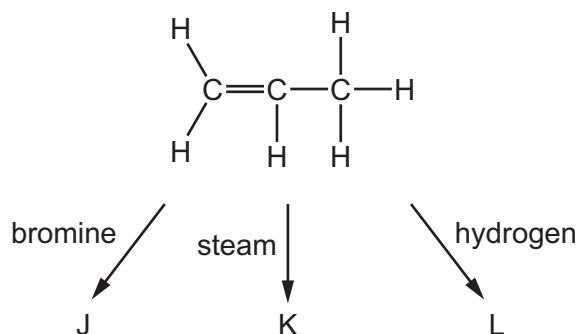
Which molecules are structural isomers of structure 1?

- A** 2 and 4      **B** 2 only      **C** 3 and 4      **D** 3 only

36 Which chemical equation for the substitution of an alkane with chlorine is correct?

- A**  $C_3H_8 + Cl_2 \rightarrow C_3H_7Cl + HCl$   
**B**  $C_3H_6 + Cl_2 \rightarrow C_3H_6Cl_2$   
**C**  $C_3H_8 + Cl_2 \rightarrow C_3H_6Cl_2 + H_2$   
**D**  $C_3H_6 + Cl_2 \rightarrow C_3H_5Cl + HCl$

37 Propene is an alkene that reacts with bromine, steam and hydrogen as shown.



What are the products of these reactions?

	J	K	L
<b>A</b>	bromopropane	propanol	butane
<b>B</b>	dibromopropane	propanoic acid	propane
<b>C</b>	dibromopropane	propanol	propane
<b>D</b>	bromopropane	propanoic acid	butane

38 The flow chart shows the preparation of ethanol and some important chemistry of ethanol.



What are X, Y and Z?

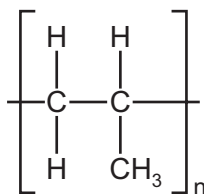
	X	Y	Z
<b>A</b>	yeast	combustion	oxygen
<b>B</b>	glucose	combustion	steam
<b>C</b>	glucose	polymerisation	water
<b>D</b>	yeast	fermentation	glucose

39 Which statements about aqueous ethanoic acid are correct?

- 1 It is an alkane.
- 2 It reacts with sodium carbonate to form carbon dioxide.
- 3 It changes the colour of litmus solution from blue to red.
- 4 It is a hydrocarbon.

**A** 1 and 2                      **B** 1 and 4                      **C** 2 and 3                      **D** 3 and 4

40 The structure of a polymer is shown.



Which monomer forms this polymer?

- A ethane
- B ethene
- C propane
- D propene

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The Periodic Table of Elements

		Group															
I	II	III	IV	V	VI	VII	VIII										
3 <b>Li</b> lithium 7	4 <b>Be</b> beryllium 9	<div style="border: 1px solid black; padding: 5px; width: fit-content; margin: auto;"> <b>Key</b>                      atomic number                      atomic symbol                      name                      relative atomic mass                 </div>										2 <b>He</b> helium 4					
11 <b>Na</b> sodium 23	12 <b>Mg</b> magnesium 24											5 <b>B</b> boron 11	6 <b>C</b> carbon 12	7 <b>N</b> nitrogen 14	8 <b>O</b> oxygen 16	9 <b>F</b> fluorine 19	10 <b>Ne</b> neon 20
19 <b>K</b> potassium 39	20 <b>Ca</b> calcium 40	21 <b>Sc</b> scandium 45	22 <b>Ti</b> titanium 48	23 <b>V</b> vanadium 51	24 <b>Cr</b> chromium 52	25 <b>Mn</b> manganese 55	26 <b>Fe</b> iron 56	27 <b>Co</b> cobalt 59	28 <b>Ni</b> nickel 59	29 <b>Cu</b> copper 64	30 <b>Zn</b> zinc 65	31 <b>Ga</b> gallium 70	32 <b>Ge</b> germanium 73	33 <b>As</b> arsenic 75	34 <b>Se</b> selenium 79	35 <b>Br</b> bromine 80	36 <b>Kr</b> krypton 84
37 <b>Rb</b> rubidium 85	38 <b>Sr</b> strontium 88	39 <b>Y</b> yttrium 89	40 <b>Zr</b> zirconium 91	41 <b>Nb</b> niobium 93	42 <b>Mo</b> molybdenum 96	43 <b>Tc</b> technetium —	44 <b>Ru</b> ruthenium 101	45 <b>Rh</b> rhodium 103	46 <b>Pd</b> palladium 106	47 <b>Ag</b> silver 108	48 <b>Cd</b> cadmium 112	49 <b>In</b> indium 115	50 <b>Sn</b> tin 119	51 <b>Sb</b> antimony 122	52 <b>Te</b> tellurium 128	53 <b>I</b> iodine 127	54 <b>Xe</b> xenon 131
55 <b>Cs</b> caesium 133	56 <b>Ba</b> barium 137	57–71 lanthanoids	72 <b>Hf</b> hafnium 178	73 <b>Ta</b> tantalum 181	74 <b>W</b> tungsten 184	75 <b>Re</b> rhenium 186	76 <b>Os</b> osmium 190	77 <b>Ir</b> iridium 192	78 <b>Pt</b> platinum 195	79 <b>Au</b> gold 197	80 <b>Hg</b> mercury 201	81 <b>Tl</b> thallium 204	82 <b>Pb</b> lead 207	83 <b>Bi</b> bismuth 209	84 <b>Po</b> polonium —	85 <b>At</b> astatine —	86 <b>Rn</b> radon —
87 <b>Fr</b> francium —	88 <b>Ra</b> radium —	89–103 actinoids	104 <b>Rf</b> rutherfordium —	105 <b>Db</b> dubnium —	106 <b>Sg</b> seaborgium —	107 <b>Bh</b> bohrium —	108 <b>Hs</b> hassium —	109 <b>Mt</b> meitnerium —	110 <b>Ds</b> darmstadtium —	111 <b>Rg</b> roentgenium —	112 <b>Cn</b> copernicium —	114 <b>Fl</b> flerovium —	116 <b>Lv</b> livermorium —	—	—	—	—

lanthanoids	57 <b>La</b> lanthanum 139	58 <b>Ce</b> cerium 140	59 <b>Pr</b> praseodymium 141	60 <b>Nd</b> neodymium 144	61 <b>Pm</b> promethium —	62 <b>Sm</b> samarium 150	63 <b>Eu</b> europium 152	64 <b>Gd</b> gadolinium 157	65 <b>Tb</b> terbium 159	66 <b>Dy</b> dysprosium 163	67 <b>Ho</b> holmium 165	68 <b>Er</b> erbium 167	69 <b>Tm</b> thulium 169	70 <b>Yb</b> ytterbium 173	71 <b>Lu</b> lutetium 175
actinoids	89 <b>Ac</b> actinium —	90 <b>Th</b> thorium 232	91 <b>Pa</b> protactinium 231	92 <b>U</b> uranium 238	93 <b>Np</b> neptunium —	94 <b>Pu</b> plutonium —	95 <b>Am</b> americium —	96 <b>Cm</b> curium —	97 <b>Bk</b> berkelium —	98 <b>Cf</b> californium —	99 <b>Es</b> einsteinium —	100 <b>Fm</b> fermium —	101 <b>Md</b> mendelevium —	102 <b>No</b> nobelium —	103 <b>Lr</b> lawrencium —

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).