



Pearson

## **Mark Scheme (Results)**

October 2017

Pearson Edexcel International Advanced  
Level In Chemistry (WCH06) Paper 01  
Chemistry Laboratory Skills II

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## General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.
- Mark schemes will indicate within the table where, and which strands of QWC, are being assessed. The strands are as follows:
  - i) ensure that text is legible and that spelling, punctuation and grammar are accurate so that meaning is clear
  - ii) select and use a form and style of writing appropriate to purpose and to complex subject matter
  - iii) organise information clearly and coherently, using specialist vocabulary when appropriate

## Using the Mark Scheme

Examiners should look for qualities to reward rather than faults to penalise. This does NOT mean giving credit for incorrect or inadequate answers, but it does mean allowing candidates to be rewarded for answers showing correct application of principles and knowledge. Examiners should therefore read carefully and consider every response: even if it is not what is expected it may be worthy of credit.

The mark scheme gives examiners:

- an idea of the types of response expected
- how individual marks are to be awarded
- the total mark for each question
- examples of responses that should NOT receive credit.

/ means that the responses are alternatives and either answer should receive full credit.

( ) means that a phrase/word is not essential for the award of the mark, but helps the examiner to get the sense of the expected answer.

Phrases/words in **bold** indicate that the meaning of the phrase or the actual word is **essential** to the answer.

ecf/TE/cq (error carried forward) means that a wrong answer given in an earlier part of a question is used correctly in answer to a later part of the same question.

Candidates must make their meaning clear to the examiner to gain the mark. Make sure that the answer makes sense. Do not give credit for correct words/phrases which are put together in a meaningless manner. Answers must be in the correct context.

### Quality of Written Communication

Questions which involve the writing of continuous prose will expect candidates to:

- write legibly, with accurate use of spelling, grammar and punctuation in order to make the meaning clear
- select and use a form and style of writing appropriate to purpose and to complex subject matter
- organise information clearly and coherently, using specialist vocabulary when appropriate.

Full marks will be awarded if the candidate has demonstrated the above abilities. Questions where QWC is likely to be particularly important are indicated (QWC) in the mark scheme, but this does not preclude others.

Question Number	Acceptable Answers	Reject	Mark
<b>1(a)</b>	<p>Any <b>two</b> from:</p> <p><math>\text{Fe}^{2+}</math> / <math>[\text{Fe}(\text{H}_2\text{O})_6]^{2+}</math> <b>(1)</b></p> <p><math>\text{Ni}^{2+}</math> / <math>[\text{Ni}(\text{H}_2\text{O})_6]^{2+}</math> <b>(1)</b></p> <p><math>\text{Cr}^{3+}</math> <b>(1)</b></p> <p>ALLOW (1) for  <math>[\text{Cr}(\text{H}_2\text{O})_6]^{3+}</math> /  <math>\text{V}^{3+}</math> / <math>[\text{V}(\text{H}_2\text{O})_6]^{3+}</math> /  <math>\text{Cu}^{2+}</math></p> <p>IGNORE  names / state symbols / missing square  brackets</p> <p>If no other mark is awarded:  ALLOW  (1) for two names with oxidation  numbers</p>	<p><math>\text{Fe}^{3+}</math></p> <p><math>\text{Cr}^{2+}</math></p>	<b>(2)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>1(b)(i)</b>	<p><math>\text{Cr}^{3+}</math> / <math>[\text{Cr}(\text{H}_2\text{O})_6]^{3+}</math></p> <p>IGNORE  names / state symbols/ missing square  brackets</p>		<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>1(b)(ii)</b>	<p><math>[\text{Cr}(\text{OH})_6]^{3-}</math></p> <p>ALLOW  <math>[\text{Cr}(\text{H}_2\text{O})_2(\text{OH})_4]^-</math> / <math>[\text{Cr}(\text{OH})_4]^-</math> / <math>\text{CrO}_2^-</math></p> <p>IGNORE  name / state symbol</p>	$\text{OH}^-$	<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>1(c)(i)</b>	Chromate <b>and</b> (VI) / VI / +6 / 6+  ALLOW CrO <sub>4</sub> <sup>2-</sup> <b>and</b> (VI) / VI / +6 / 6+	Incorrect formula e.g. Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>	<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>1(c)(ii)</b>	Gas: ammonia / NH <sub>3</sub> <b>(1)</b>  Ionic equation: NH <sub>4</sub> <sup>+</sup> + OH <sup>-</sup> → NH <sub>3</sub> + H <sub>2</sub> O <b>(1)</b>  IGNORE state symbols, even if incorrect / non-ionic equation	ammonium / NH <sub>4</sub> / NH <sub>4</sub> <sup>+</sup>	<b>(2)</b>

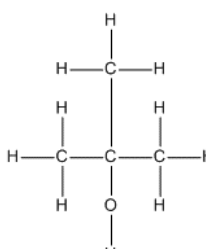
Question Number	Acceptable Answers	Reject	Mark
<b>1(d)</b>	Barium sulfate / BaSO <sub>4</sub>		<b>(1)</b>




Question Number	Acceptable Answers	Reject	Mark
<b>1(e)</b>	NH <sub>4</sub> Cr(SO <sub>4</sub> ) <sub>2</sub> / (NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub> .Cr <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>  ALLOW Correct neutral formulae with any Cr <sup>3+</sup> : NH <sub>4</sub> <sup>+</sup> ratio e.g. Cr(NH <sub>4</sub> ) <sub>3</sub> (SO <sub>4</sub> ) <sub>3</sub> / (NH <sub>4</sub> ) <sub>2</sub> Cr <sub>2</sub> (SO <sub>4</sub> ) <sub>4</sub>  ALLOW species in any order  ALLOW species with charges e.g. NH <sub>4</sub> <sup>+</sup> Cr <sup>3+</sup> (SO <sub>4</sub> <sup>2-</sup> ) <sub>2</sub>  IGNORE Missing dot / water of crystallisation  TE on (b)(i) (c)(ii) and (d) e.g. (NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub> .FeSO <sub>4</sub>	Species with charges giving a net charge	<b>(1)</b>

**(Total for Question 1 = 9 marks)**

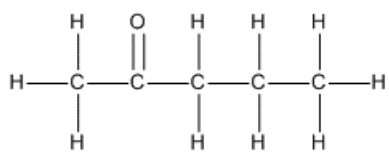
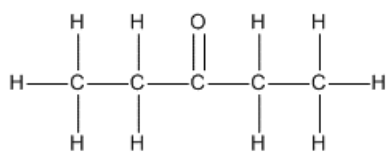
Question Number	Acceptable Answers	Reject	Mark
<b>2(a)(i)</b>	(Gas) hydrogen chloride / HCl / HCl(g)  ALLOW HCl(aq) / hydrochloric acid <b>(1)</b>  (Group) OH / -OH / alcohol / hydroxy / hydroxyl <b>(1)</b>	OH <sup>-</sup> / hydroxide / carboxylic acid	<b>(2)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>2(a)(ii)</b>	Tertiary / 3° alcohol  ALLOW Not primary <b>and</b> not secondary alcohol  IGNORE <b>D</b> cannot be oxidised / not aldehyde / just '-OH'	Carboxylic acid	<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>2(b)</b>	 <p>ALLOW OH / CH<sub>3</sub></p> <p>IGNORE connectivity to vertical OH but penalise OH-C on left</p> <p>ALLOW displayed formula of propanoic acid</p>		<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>2(c)(i)</b>	Penalise methyl ketone for ketone once only in (c)(i) and (c)(ii)  Aldehyde / CHO /  /  or ketone /  $R_2C=O$ /  Both aldehyde and ketone needed for the mark  ALLOW Carbonyl  IGNORE Just C=O	Carboxyl	<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>2(c)(ii)</b>	Ketone		<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>2(d)(i)</b>	    ALLOW Isomers in either order Structural / skeletal / displayed formulae or any combination of these  TE if aldehyde identified in (c)(ii) for (1) mark	Incorrect names in addition to formulae once only  <b>(1)</b> Minor errors once only e.g. H missing / bond missing  <b>(1)</b> Pentavalent carbons in each structure	<b>(2)</b>



Question Number	Acceptable Answers	Reject	Mark
<b>2(d)(ii)</b>	<p><b>Reagents</b> – standalone mark</p> <p>Iodine / I<sub>2</sub> (in potassium iodide / KI) / KI<sub>3</sub>  <b>and</b>  sodium hydroxide / NaOH (solutions)  OR  Potassium iodide / KI  <b>and</b>  sodium chlorate(I) / NaOCl (solutions)</p> <p>ALLOW  Iodine and alkali <b>(1)</b></p> <p><b>Observations</b> – conditional on at least one ketone in (d)(i)</p> <p>(Observation for pentan-2-one)  (Pale) yellow precipitate /ppt /ppte / solid  (with antiseptic smell) <b>(1)</b></p> <p>IGNORE yellow colour</p> <p>(Observation for pentan-3-one)  No change / no precipitate /no ppt/ no ppte  / no solid</p> <p>ALLOW  No reaction  No change in colour / remains colourless /  remains (pale) yellow <b>(1)</b></p> <p>No TE on structures in (d)(i)</p> <p>IGNORE  References to spectroscopy</p>	<p>NH<sub>3</sub> as alkali</p> <p>White ppt</p> <p>Other colours  e.g. brown</p>	<b>(3)</b>

**(Total for Question 2 = 11 marks)**

Question Number	Acceptable Answers	Reject	Mark
<b>3(a)</b>	<p>Correct answer, with or without working, scores (3)</p> <p><b>First mark</b> - calculating moles of O<sub>2</sub>  Mol O<sub>2</sub> = 20/24000 or 0.02/24  = 0.00083333 / 8.3333 x 10<sup>-4</sup>  ALLOW 1/1200  OR  Mol O<sub>2</sub> = 20/24  = 0.833 <b>(1)</b></p> <p><b>Second mark</b> - using mole ratio  Mol H<sub>2</sub>O<sub>2</sub> (in 1 cm<sup>3</sup>)  = 2 x 0.00083333  = 0.00167 / 1.67 x 10<sup>-3</sup>  ALLOW 1/600  OR  Mol H<sub>2</sub>O<sub>2</sub> (in 1 dm<sup>3</sup>)  = 2 x 0.833  = 1.67  TE on mol O<sub>2</sub> <b>(1)</b></p> <p><b>Third mark</b> – calculating concentration of H<sub>2</sub>O<sub>2</sub>  This is dependent on a mole ratio being used in the calculation</p> <p>Concentration H<sub>2</sub>O<sub>2</sub>  = 0.00167 x 1000  = 1.67 (mol dm<sup>-3</sup>)  OR  Recognition that there are 1.67 mol H<sub>2</sub>O<sub>2</sub> in 1 dm<sup>3</sup> so concentration = 1.67 mol dm<sup>-3</sup>  TE on mol H<sub>2</sub>O<sub>2</sub></p> <p>ALLOW 5/3 and 1<sup>2</sup>/3 <b>(1)</b></p> <p>IGNORE SF except 1 SF</p> <p>Do not penalise final answer from correct rounding of intermediate answers e.g. 1.6 / 1.66 (mol dm<sup>-3</sup>) from 0.83 / 0.000833</p>		<b>(3)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>3(b)(i)</b>	(Rinse it with distilled water followed by) rinse it with the hydrogen peroxide / solution (it will be filled with)  ALLOW rinse with water and hydrogen peroxide  IGNORE just clean / rinse / dry	Rinse with hydrogen peroxide and water  Rinse with hydrogen peroxide then water	<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>3(b)(ii)</b>	Shake / invert the solution  ALLOW any indication of mixing e.g.stir		<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>3(b)(iii)</b>	The nozzle / jet of the burette is full / has no air bubbles  ALLOW  The part below the tap is full (of solution) OR The gap is full (of solution) / the bottom part is full (of solution) OR There are no (air) bubbles / check for air gaps OR The funnel has been removed  IGNORE References to reading from the bottom of the meniscus / reading at eye level Check that the tap is closed		<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>3(b)(iv)</b>	(25 cm <sup>3</sup> ) measuring cylinder  ALLOW Beaker with 25 cm <sup>3</sup> graduation mark / 50 cm <sup>3</sup> beaker / 100 cm <sup>3</sup> beaker / 100 cm <sup>3</sup> or 50 cm <sup>3</sup> conical flask with graduation mark / syringe  IGNORE just 'beaker' / 'conical flask'	pipette burette volumetric flask	<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>3(b)(v)</b>	(From) colourless (to permanent pale) pink  Both colours needed for the mark  IGNORE clear	Colourless to purple / mauve / red  Pink / purple / mauve to colourless	<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>3(b)(vi)</b>	Results that are within 0.1 / 0.2 (cm <sup>3</sup> of each other)  ALLOW Results that are similar <b>and</b> a specific example within 0.1 / 0.2 e.g. 25.00 and 25.20 (cm <sup>3</sup> )  Results that are within $\pm 0.1$ / $\pm 0.2$ (cm <sup>3</sup> of each other)  IGNORE Just 'results that are (almost) the same' / results that are close together / results within a stated % of each other		<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>3(b)(vii)</b>	<p>Correct answer, with or without working, scores (4)</p> <p>Mol <math>\text{MnO}_4^-</math> used  <math>= \frac{15.80 \times 0.0200}{1000}</math>  <math>= 0.000316 / 3.16 \times 10^{-4}</math> <b>(1)</b></p> <p>Mol <math>\text{H}_2\text{O}_2</math> in <math>25.0 \text{ cm}^3</math>  <math>= \frac{0.000316 \times 5}{2}</math>  <math>= 0.000790 / 7.90 \times 10^{-4}</math></p> <p>TE on mol <math>\text{MnO}_4^-</math> <b>(1)</b></p> <p>Concentration diluted <math>\text{H}_2\text{O}_2</math>  <math>= \frac{0.000790 \times 1000}{25.0}</math>  <math>= 0.0316 \text{ (mol dm}^{-3}\text{)}</math></p> <p>TE on mol <math>\text{H}_2\text{O}_2</math> in <math>25.0 \text{ cm}^3</math> <b>(1)</b></p> <p>Concentration original <math>\text{H}_2\text{O}_2</math>  <math>= \frac{0.0316 \times 250.0}{10.0}</math>  <math>= 0.790 / 0.79 \text{ (mol dm}^{-3}\text{)}</math></p> <p>TE on concentration diluted <math>\text{H}_2\text{O}_2</math> <b>(1)</b></p> <p><b>Alternative method for 3<sup>rd</sup> and 4<sup>th</sup> marks</b></p> <p>Mol <math>\text{H}_2\text{O}_2</math> in <math>250.0 \text{ cm}^3</math> / (original) <math>10.0 \text{ cm}^3</math>  <math>= 0.000790 \times 10</math> or <math>0.000790 \times</math>  <u>250.0</u></p> <p>25.0  <math>= 0.00790</math></p> <p>TE on mol <math>\text{H}_2\text{O}_2</math> in <math>25.0 \text{ cm}^3</math> <b>(1)</b></p> <p>Concentration original <math>\text{H}_2\text{O}_2</math>  <math>= 0.00790 \times \frac{1000}{10.0}</math>  <math>= 0.790 / 0.79 \text{ (mol dm}^{-3}\text{)}</math></p> <p>TE on mol <math>\text{H}_2\text{O}_2</math> in <math>250.0 \text{ cm}^3</math> / (original) <math>10.0 \text{ cm}</math> <b>(1)</b></p> <p>IGNORE SF except 1SF</p>		<b>(4)</b>

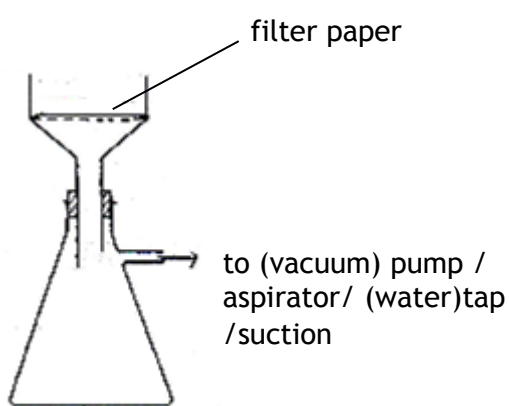
Question Number	Acceptable Answers	Reject	Mark
<b>3(c)</b>	$\frac{2 \times 0.05}{15.80} \times 100 =$  (±)0.63291 / 0.6329 / 0.633 / 0.63 / 0.6(%)	0.31646 (%)	<b>(1)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>3(d)</b>	(Some of) the hydrogen peroxide had already decomposed (slowly, to form water and oxygen)  OR (Some of) the hydrogen peroxide had formed / reacted to form oxygen / water  OR (Some of) the hydrogen peroxide had undergone disproportionation / oxidation and reduction  IGNORE Just oxidation / just reduction  Hydrogen peroxide has low boiling / melting temperature  Water formed dilutes the hydrogen peroxide	Any references to a reaction with $\text{MnO}_4^-$  Water /hydrogen peroxide evaporates  Hydrogen is formed	<b>(1)</b>

**(Total for Question 3 = 15 marks)**

Question Number	Acceptable Answers	Reject	Mark
<b>4(a)(i)</b>	<p>As a (mutual/co) solvent (for the ester and sodium hydroxide solution) OR Dissolves ester and sodium hydroxide / both substances / the mixture OR To allow the ester and sodium hydroxide (solution) to mix / dissolve / become miscible</p> <p>ALLOW Dissolves the ester</p> <p>IGNORE To allow the ester and sodium hydroxide to react / References to the product dissolving</p>	Catalyst	<b>(1)</b>

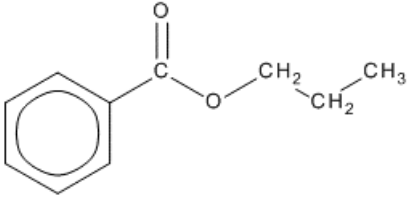
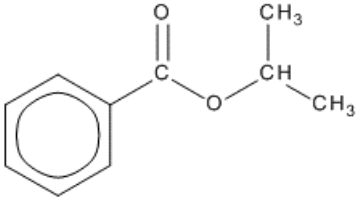
Question Number	Acceptable Answers	Reject	Mark
<b>4(a)(ii)</b>	<p><b>First mark</b> Heating to increase / speed up the rate of reaction / make the reaction faster</p> <p>ALLOW To overcome / provide the (high) activation energy / reaction is slow at room temperature <b>(1)</b></p> <p>IGNORE Just 'to provide energy for the reaction'</p> <p><b>Second mark</b> Refluxing to prevent loss of (volatile) reactants / products OR So that (volatile) reactants / products return to the flask</p> <p>ALLOW To condense the gases / vapours formed To prevent gas / vapour escaping So that reaction goes to completion <b>(1)</b></p> <p>IGNORE To condense the mixture / any reference to flammable gases</p>	Just 'to prevent evaporation'	<b>(2)</b>

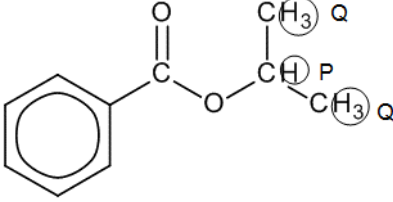
Question Number	Acceptable Answers	Reject	Mark
<b>4(a)(iii)</b>	 <p>(Top label) Filter paper</p> <p>ALLOW Sintered glass <b>(1)</b></p> <p>IGNORE Porous filter / Buchner funnel</p> <p>(Lower label) To (vacuum) pump / aspirator / (water) tap / suction</p> <p>ALLOW To vacuum (apparatus) / reduced pressure <b>(1)</b></p> <p>(Reason) Faster / speeds up (filtration) OR (Product / benzoic acid is) drier</p> <p>ALLOW Filtrate / soluble impurities / solvent is removed <b>and</b> more completely / efficiently OR Dries the benzoic acid <b>(1)</b></p>	<p>Fluted filter paper</p> <p>Waste gas / gas out / pressure out</p> <p>Reference to removing insoluble impurities</p>	<b>(3)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>4(a)(iv)</b>	<p>Dissolve/ add the benzoic acid / (impure) solid / crystals in the minimum (volume / amount) of boiling / hot water</p> <p>ALLOW use of solvent for water</p>	Incorrect solvent	<b>(1)</b>



Question Number	Acceptable Answers	Reject	Mark
<b>4(a)(v)</b>	<p>Correct answer, with or without working scores (4)</p> <p><b>First mark</b> Mass of ester/<b>X</b> = <math>3.0 \times 1.02 = 3.06</math> (g) <b>(1)</b></p> <p><b>Second mark</b> moles of ester/<b>X</b> = <math>\frac{3.06}{164} = 0.018659</math> <b>(1)</b></p> <p><b>Third mark</b> <b>EITHER</b> Theoretical mass benzoic acid = <math>0.018659 \times 122</math> = <math>2.2763</math> (g)</p> <p>TE on moles of ester/<b>X</b></p> <p><b>OR</b> Moles of benzoic acid produced = <math>1.45/122</math> = <math>0.011885</math></p> <p>TE on moles of ester/<b>X</b> <b>(1)</b></p> <p><b>Fourth mark</b> <b>EITHER</b> % yield = <math>\frac{1.45}{2.2763} \times 100 = 63.6987(\%)</math></p> <p>TE on theoretical mass benzoic acid</p> <p><b>OR</b> % yield = <math>\frac{0.011885}{0.018659} \times 100 = 63.6987(\%)</math></p> <p>TE on moles benzoic acid produced <b>(1)</b></p> <p>IGNORE SF except 1 SF</p>		<b>(4)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>4(b)</b>	 <p style="text-align: right;"><b>(1)</b></p>  <p style="text-align: right;"><b>(1)</b></p> <p>Structures can be in either order</p> <p>ALLOW displayed, skeletal, structural formulae or any combination of these</p>	Any missing H from non-skeletal formulae once only	<b>(2)</b>

Question Number	Acceptable Answers	Reject	Mark
<b>4(c)</b>	 <p><b>First mark</b> Identification of <b>X</b> as isopropylbenzoate <b>(1)</b></p> <p><b>Second mark</b> Identification of proton responsible for peak P <b>and</b> Identification of <b>both</b> sets of protons responsible for peak Q</p> <p>ALLOW carbon atoms / CH and CH<sub>3</sub> groups labelled</p> <p>ALLOW just 1 proton in each CH<sub>3</sub> group labelled <b>(1)</b></p> <p>IGNORE Missing benzoate group / circle in benzene</p>	Incorrect ester scores (0)	<b>(2)</b>

**(Total for Question 4 = 15 marks)**